

## Modern Chemistry Chapter 14 Answers

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### Modern Chemistry Chapter 14 Answers

Modern Chemistry 121 Acids and Bases CHAPTER 14 REVIEW Acids and Bases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Answer the following questions according to the Brønsted-Lowry definitions of acids and bases: \_\_\_\_\_ a.

### Modern Chemistry Chapter 14 Review Answers Acids And Bases

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Modern Chemistry Chapter 14 vocabulary. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. mesia456. Terms in this set (21) binary acid. an acid that does not contain oxygen, such as hydrofluoric acid. oxyacid. an acid that is a compound of hydrogen, oxygen, and a third element, usually a nonmetal.

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### Modern Chemistry Chapter 14 Mixed Review Answers

Modern Chemistry Chapter 14 & 15. binary acid. oxyacid. strong acid. weak acid. an acid that does not contain oxygen, such as hydrofluoric acid. an acid that is a compound of hydrogen,oxygen, and a third ele.... an acid that ionizes completely in a solvent.

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CHAPTER 14 REVIEW Acids and Bases SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. a. Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1:  $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$  stage 2:  $\text{HSO}_3^-(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{SO}_3^{2-}(\text{aq})$  b.

### 14 Acids and Bases

Modern Chemistry Chapter 14 Acids & Bases Section 1 Properties of Acids & Bases Aqueous solutions of acids taste sour. Acids change the color of acid-base indicators. Some acids react with active metals and release hydrogen gas. Acids react with bases to produce salts and water. Acids conduct electricity (electrolytes).

### Modern Chemistry Chapter 14

CHAPTER 14 REVIEW . Acids and Bases. SHORT ANSWER Answer the following questions in the space provided. 1. a. Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1:  $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$  b. Which stage of ionization usually produces more ions? Explain your answer.

### CHAPTER 14 REVIEW Acids and Bases

Chemistry (4th Edition) Burdge, Julia Publisher McGraw-Hill Publishing Company ISBN 978-0-07802-152-7

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### [MOBI] Modern Chemistry Chapter 14 Section 1 Answers

Modern Chemistry Chapter 14. boiling-point elevation. colligative properties. dissociation. freezing-point depression. the difference between the boiling point of a pure... solvent and.... properties that depend on the concentration of solute... particle.... the separation of ions that occurs when an ionic compound... diss....

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CHAPTER 14 REVIEW Ions in Aqueous Solutions and Colligative Properties SECTION 14-1 SHORT ANSWER Answer the following questions in the space provided. 1. Use the guidelines in Table 14-1 on page 427 of the text to predict the solubility of the following compounds in water: a. magnesium nitrate b. barium sulfate c. calcium carbonate d. ammonium phosphate

### 14 Ions in Aqueous Solutions and Colligative Properties

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CHAPTER 14 REVIEW Acids and Bases SECTION 2 SHORT ANSWER Answer the following questions in the space provided 1 a Explain your answers BCl 3 is the Lewis acid because it accepts an electron pair in forming a covalent bond It is not a Brønsted-Lowry acid, because it is not donating a proton 122 ACIDS AND BASES MODERN CHEMISTRY Study ...

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CHAPTER 16 REVIEW Reaction Energy SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. For elements in their standard state, the value of  $\Delta H_f^\circ$  is . 2. The formation and decomposition of water can be represented by the following thermochemical equations:  $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{g})$  241.8 kJ/mol  $\text{H}_2\text{O}(\text{l})$  241.8 kJ/mol ...

### 16 Reaction Energy

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