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Both ions have a  $-2$  charge, and each ion includes both sulfur and oxygen atoms. Sulfate has four oxygen atoms; sulfite only has three. 9. The "chlorate" polyatomic ion has a charge of  $-1$  and is composed of one chlorine atom (the central atom) and three oxygen atoms. a. Draw a model of a chlorate ion. 124  
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6. The dotted line around each polyatomic ion in Model 1 shows that the group of atoms has a charge. The charge is not on any one atom, but rather on the group of atoms as a whole. Based on your knowledge of monatomic ions, propose an explanation for the net charge on a polyatomic ion. Your answer should include a discussion of subatomic ...

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Based on your knowledge of monatomic ions, propose an explanation for the net charge on a polyatomic ion. Your answer should include a discussion of subatomic particles. Somewhere in the ion there are either extra (in the case of an anion) or fewer electrons (in the case of a cation) compared to the total number of protons in the atoms involved. 7.

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3. In a sample of solid  $\text{Ba}(\text{NO}_3)_2$ , the ratio of barium ions to nitrate ions is 1) ammonia nitrite 2) ammonium nitrite 3) ammonia

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nitrate 4) ammonium nitrate 4. What is the correct name of the compound with the formula  $\text{NH}_4\text{NO}_2$ ? 1) hydroxide 2) oxalate 3) oxide 4) peroxide 5. What is the name of the polyatomic ion in the compound  $\text{Na}_2\text{O}_2$ ? 1)  $\text{NaClO}$  2)  $\text{Na}_2\text{O}$  3 ...

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